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Example

\ (\PageIndex{1}) :

Electronegativity  
and Bond Polarity.

Bond polarities  
play an important  
role in

determining the  
structure of  
proteins. Using  
the

electronegativity  
values in Figure

\ (\PageIndex{1}) ,  
arrange the

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following covalent bonds—all commonly found in amino acids—in order of increasing polarity. Then designate the positive and negative atoms using the symbols  $\delta^+$  and  $\delta^-$ :

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problem  
demonstrates how  
to use  
electronegativity  
to determine bond  
polarity and  
whether or not a  
bond is more  
covalent or more  
ionic.

Electronegativity  
- Dipole, VSEPR  
Shape, Symmetry,  
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is an additional  
energy that comes  
from ionic  
factors, i.e.  
polar character of  
the bond.

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In your paragraph,  
describe why  
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values and not  
absolute  
electronegativity  
values when  
determining the  
polarity of a

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bond. Also explain  
how a molecule  
with polar bonds  
can be non-polar  
overall. Answer

Key: Making  
Connections

between

Electronegativity,  
Molecular Shape,  
and Polarity .

Question 3:

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A polar molecules  
is a molecule with  
a difference in  
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two ends. 2.

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values in Figure 3, arrange the following covalent bonds—all commonly found in amino acids—in order of increasing polarity. Then designate the positive and negative atoms using the symbols  $\delta^+$  and  $\delta^-$ : C-H, C-N, C-O, N-H, O-H, S-H ...

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Relative Strengths  
Of Acids And Bases

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Atom 1 Atom 2  
Electronegativity  
Difference ( $\Delta EN$ )  
Bond Type  
(Nonpolar Covalent  
(NPC), Moderately  
Polar Covalent  
(MPC), Very Polar

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And  
Covalent (VPC), or  
Ionic (I) Arsenic

Sulfur Cobalt

Bromine Germanium

Selenium Silicon

Fluorine Potassium

Nitrogen Nickel

Oxygen Barium Tin

Hydrogen Oxygen

Calcium Sulfur

Iron Carbon Part B

Directions: Draw

the Lewis Dot

Structure for each

compound below.



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Then ...

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A study of  
electronegativity,  
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molecular polarity  
Introduction: In

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this atomic-level  
simulation, you  
will investigate  
how atoms'

electronegativity  
value affects the  
bonds they produce  
When two atoms  
bond, a pair of  
electrons is  
shared

.